|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| **Name** | **Symbol** | **Group Number** | **Valence Electrons** | **Ion Formed** |
| Oxygen |  |  |  |  |
|  | Mg |  |  |  |
| Sodium |  |  |  |  |
|  | F |  |  |  |
| Argon |  |  |  |  |
|  | Al |  |  |  |
| Nitrogen |  |  |  |  |
|  | Br |  |  |  |
|  | K |  |  |  |
| Neon |  |  |  |  |
| Calcium |  |  |  |  |
|  | B |  |  |  |
|  | Sr |  |  |  |
| Sulfur |  |  |  |  |

Name: Period:

Ion Formation Practice

Directions: Complete the table below by using the formula of each compound to identify the elements that each compound contains and the number of atoms of each of these elements in a unit of the compound.

|  |  |  |  |
| --- | --- | --- | --- |
| Formula | Element 1 | Element 2 | Element 3 |
| CO2 |  |  | --------------------- |
| HClO4 |  |  |  |
| Pb(NO3)2 |  |  |  |
| H2(SO4)2 |  |  |  |
| C6H6 |  |  | ---------------------- |
| C6H8O12 |  |  |  |

Directions: Draw Electron Dot Diagrams for the elements listed.

1. Carbon e. Argon
2. Sulfur f. Phosphorus
3. Chlorine g. Magnesium
4. Helium h. Aluminum